

### Balancing Chemical Equations Practice

1.  $\text{H}_2 + \text{Cl}_2 \rightarrow \text{HCl}$
2.  $\text{HgO} \rightarrow \text{Hg} + \text{O}_2$
3.  $\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$
4.  $\text{H}_2\text{O} \rightarrow \text{H}_2 + \text{O}_2$
5.  $\text{Al} + \text{Br}_2 \rightarrow \text{AlBr}_3$
6.  $\text{Hf} + \text{N}_2 \rightarrow \text{Hf}_3\text{N}_4$
7.  $\text{Cr}_2\text{O}_3 \rightarrow \text{Cr} + \text{O}_2$
8.  $\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$
9.  $\text{N}_2\text{H}_4 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{N}_2$
10.  $\text{F}_2 + \text{H}_2\text{O} \rightarrow \text{HF} + \text{O}_3$
11.  $\text{H}_3\text{PO}_4 + \text{NaOH} \rightarrow \text{Na}_3\text{PO}_4 + \text{H}_2\text{O}$
12.  $\text{Cu} + \text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + \text{Ag}$
13.  $\text{Ca}(\text{NO}_3)_2 + \text{KOH} \rightarrow \text{Ca}(\text{OH})_2 + \text{KNO}_3$
14.  $\text{Al} + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{Al}(\text{NO}_3)_3 + \text{Pb}$
15. sodium reacts with water to produce sodium hydroxide and hydrogen gas
16. chlorine and sodium bromide react to form bromine and sodium chloride
17. copper combines with sulfur to form copper(I) sulfide