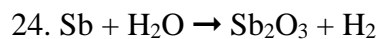
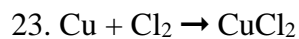
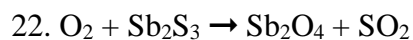
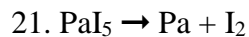
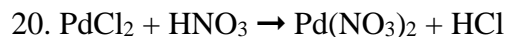
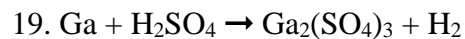
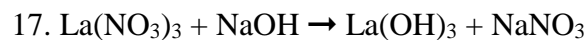
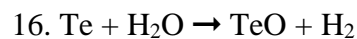
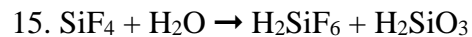
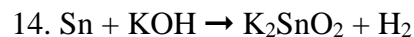
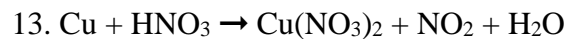
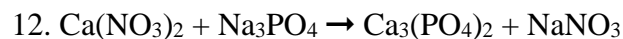
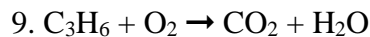
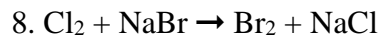
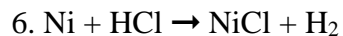
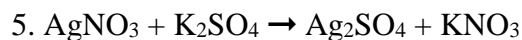
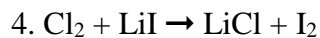
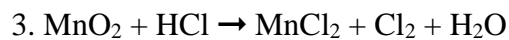
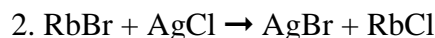
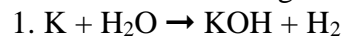
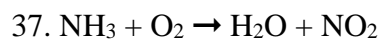
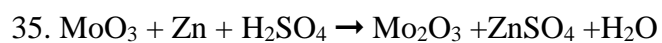
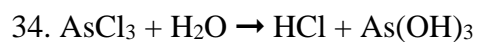
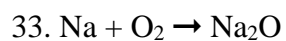
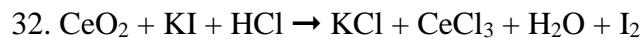
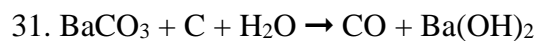
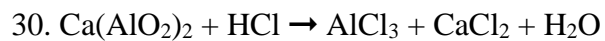
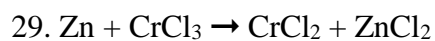
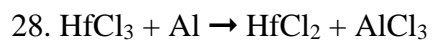
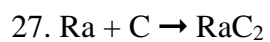
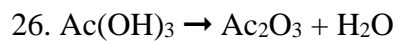


Balancing Chemical Equations Practice

Balance the following reactions:





Balance each of the following:

1. copper(II) carbonate decomposes to copper(II) oxide and carbon dioxide
2. ammonium nitrite decomposes to nitrogen gas and water
3. silver nitrate reacts with sulfuric acid to produce silver sulfate and nitric acid
4. sulfuric acid decomposes to water and sulfur trioxide
5. calcium carbonate reacts with hydrochloric acid to produce calcium chloride, water, and carbon dioxide
6. ammonium nitrate decomposes to water and dinitrogen monoxide
7. aluminum and iron(III)oxide produce iron and aluminum oxide
8. aluminum and sulfuric acid form hydrogen gas and aluminum sulfate $(2\text{Al} + 3\text{H}_2\text{SO}_4(\text{aq}) \rightarrow 3\text{H}_2 + \text{Al}_2(\text{SO}_4)_3)$
9. calcium phosphate reacts with sulfuric acid to make phosphoric acid and calcium sulphate.
 $(\text{Ca}_3(\text{PO}_4)_2 + \text{H}_2\text{SO}_4 \rightarrow \text{H}_3\text{PO}_4 + \text{CaSO}_4)$
10. copper metal reacts with sulfuric acid to form copper(II) sulfate, sulfur dioxide and water. $(\text{Cu} + 2\text{H}_2\text{SO}_4 \rightarrow \text{CuSO}_4 + \text{SO}_2 + \text{H}_2\text{O})$
11. aluminum and water react to produce aluminum hydroxide and hydrogen gas.
12. hydrochloric acid and magnesium react to form magnesium chloride and hydrogen gas.
13. calcium carbonate breaks down into elements
14. bromine and potassium iodide react to produce iodine and potassium bromide
15. sulfuric acid and calcium chlorite form hydrogen chlorite and calcium sulfate
16. copper (II) sulfide and oxygen react to produce copper(II) oxide and sulfur
17. silver oxide decomposes into its elements